

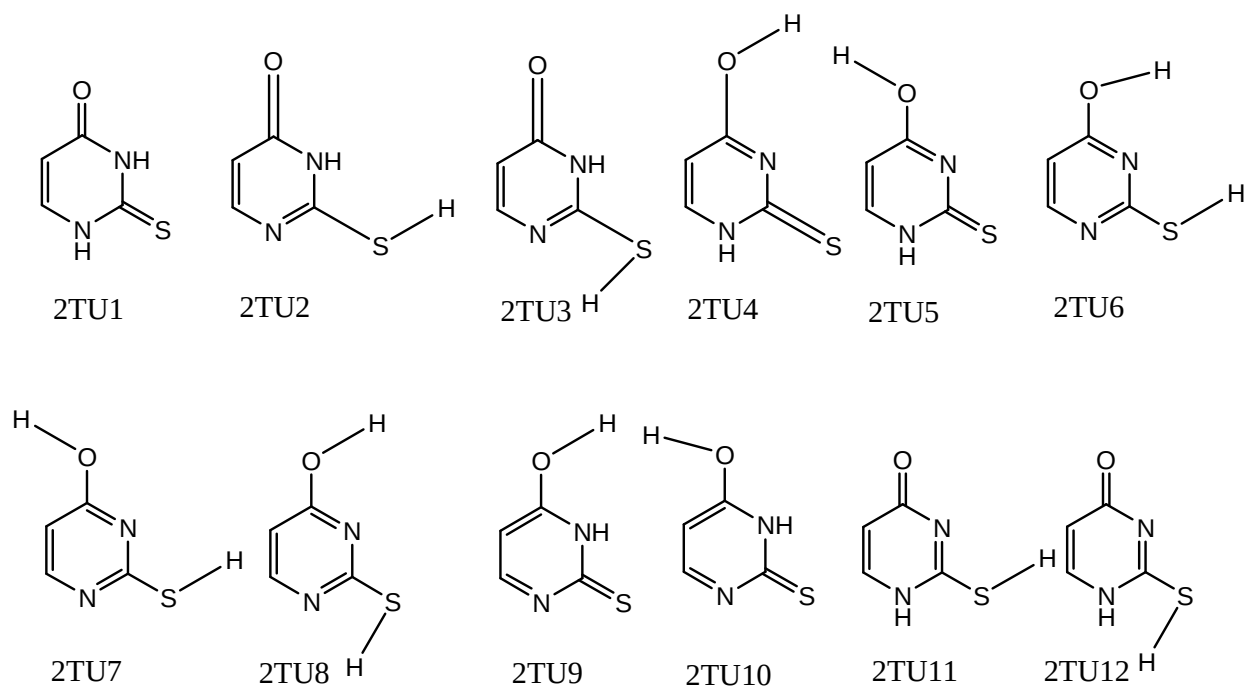


Fig. 1. Tautomers of 2-thiouracil

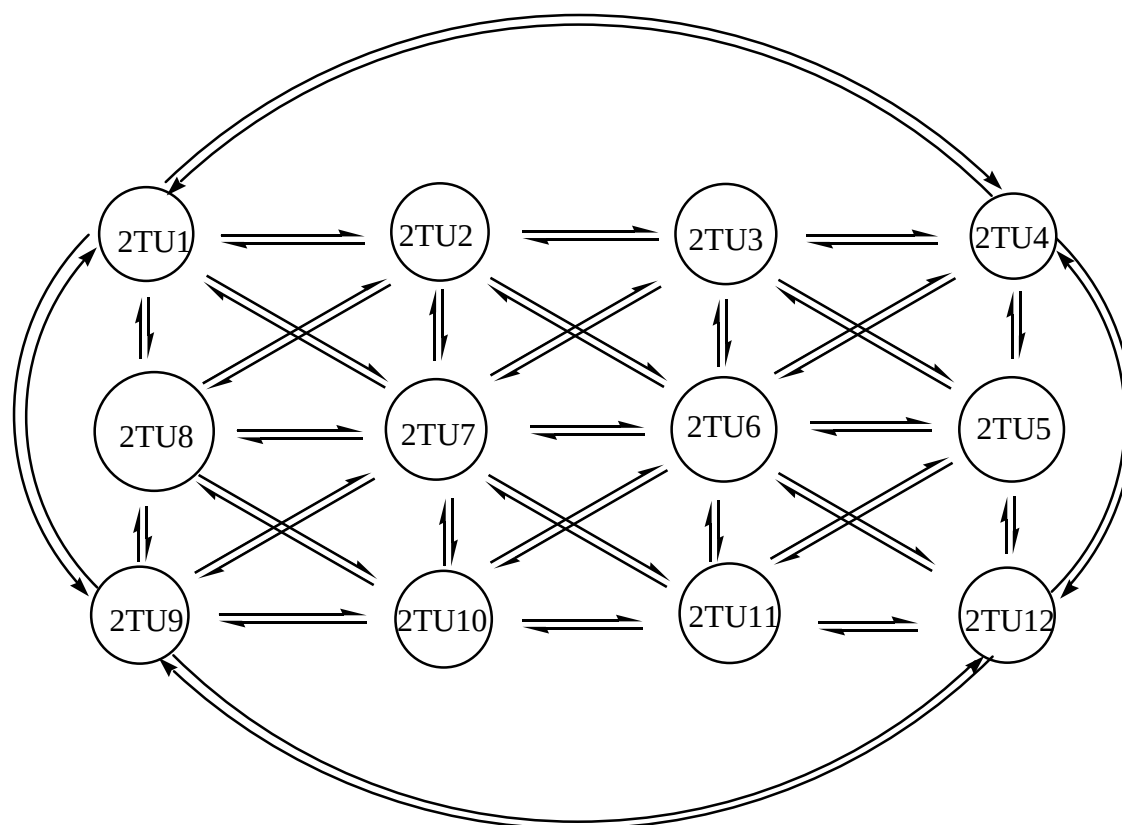


Fig. 3. Diagram of interconversion of 2-thiouracil

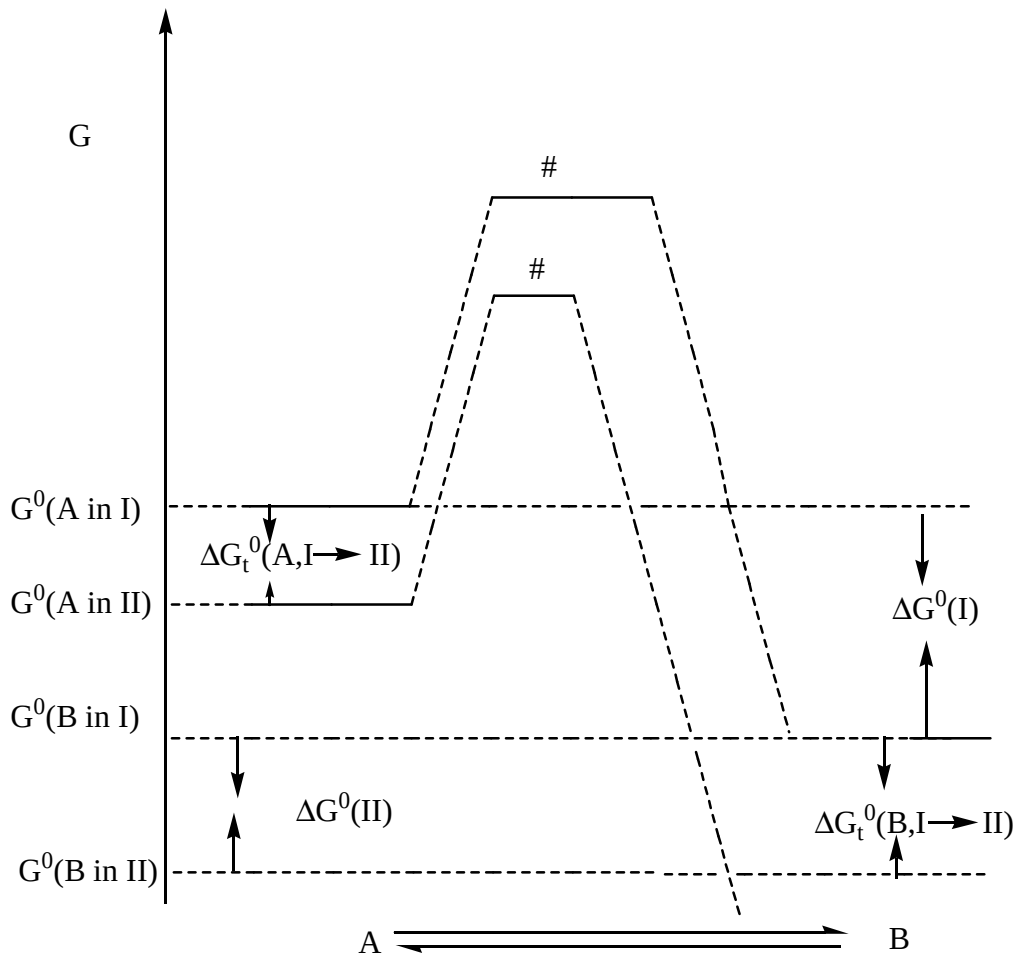


Fig. 2. One-dimensional Gibbs energy diagram for an equilibrium reaction $A \leftrightarrow B$ in the solvents I and II. Ordinate: standard molar Gibbs energies of the reactants A and B in solvents I and II; Abscissa: not defined. $\Delta G^0(\text{I})$ and $\Delta G^0(\text{II})$: standard molar Gibbs energies of reaction in solvents I and II, respectively; $\Delta G_t^0(\text{A, I} \rightarrow \text{II})$ and $\Delta G_t^0(\text{B, I} \rightarrow \text{II})$: standard molar Gibbs energies of transfer of the solutes A and B from solvent I to solvent II, respectively [$\Delta G_t^0(\text{A, I} \rightarrow \text{II}) = G^0(\text{A in I}) - G^0(\text{A in II})$, and [$\Delta G_t^0(\text{B, I} \rightarrow \text{II}) = G^0(\text{B in I}) - G^0(\text{B in II})$], $\#$ = transition state.