

MONITORING AND EVALUTION OF EXOTHERMS AS TROUBLE-SHOOTING PROCESSES IN POLYURETHANE FOAM MANUFACTURE

ABSTRACT

Monitoring and evalution of exotherms as temperature rises were investigated using various foam formulations from two types of polyols with variations in blow indices as volume of water added. It was observed that exotherms increase linearly with blow index. The results also showed that density variations are in inverse relationship with the exotherms, foam heights and compression sets. Furthermore, it was similarly observed that as the blow index increases, ΔT_1 and ΔT_2 increase for both conventional (CPO) and polymer (PPO) polyols far above the acceptable temperature range of 22-26⁰C for foam manufacture with consequences on physical properties.

Keywords- conventional polyols, polymeric polyol, exotherms, density, compression set

INTRODUCTION

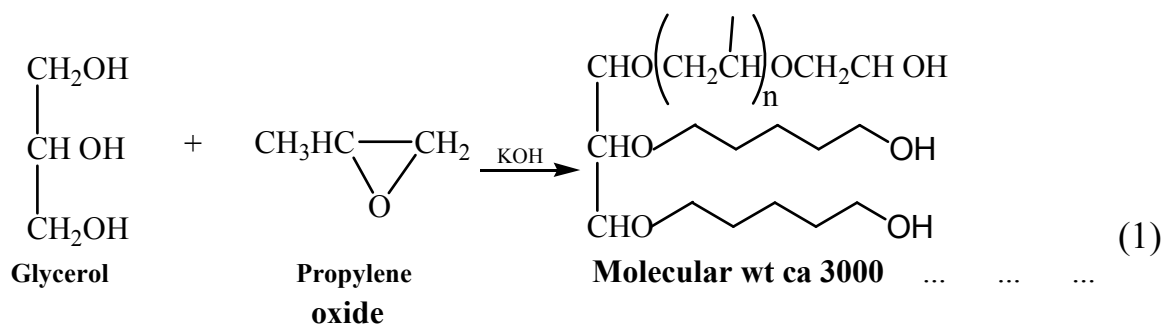
Polvinylchloride (PVC), polystyrene (PS), polyurethane (PU), acrylonitrile butadiene-styrene (ABS), and cellulose acetate are some of the plastics that can be commercially produced in cellular or expanded forms as foams [1]. Foamed plastics comprise a versatile new family of low density, cellular materials which are made in a variety of processes resembling the making of bread whereby a cake rises with the decomposition of the baking powder [2]. Foams, generally have low densities with high modulus to density and high strength to weight ratios and are good heat insulators due to the low conductivity of a gas (air) contained in the matrix [3-4].

Expanded plastics (plastics foams or cellular plastics) have been reported in several literature materials. If a gas is introduced chemically or mechanically during polymerization, a cell-like structure is left of the polymer material, giving it the characteristic foam or cellular appearance [2].

Foamed plastics can be classified based on several schemes, one of which is the nature of the foam cells [4-5]. A closed cell foam has each individual cell, more or less spherical in shape, and is completely closed in a wall of plastics. Closed cell foams are usually produced in processes where some pressure is maintained and the gas is entrapped during cell formation stage, for instance, in the reaction of isocyanate with water in the manufacture of urethane foams [6]. On the other hand, open cell foams have individual cells cross-linked as sponges as a result of free expansion during cell formation. Open cell foams are more suitable for cushioning application because they allow for compression and consequently air flow between cells [7].

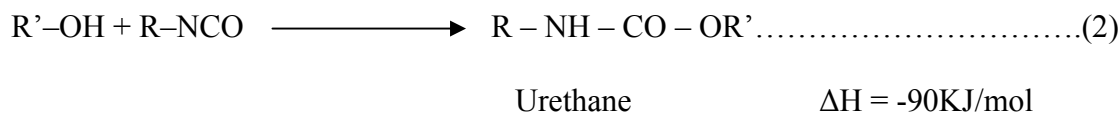
Polyurethane foams are formed by reacting a monomer containing at least two isocyanate functional groups $R-(N=C=O)$ with another monomer containing at least two or more hydroxyl groups in the presence of a catalyst system and other additives. These may be aliphatic or aromatic isocyanates and polyols (polyether or polyester resins) [7-8]. Flexible polyurethane foams are produced from linear or high molecular weight polyols that contain few hydroxyl groups per molecule, which on reaction with a isocyanate, give structures with low degree of cross-links [9]. On the other hand, polyols of high functionalities react with isocyanates to produce rigid foams with high degree of cross linking. These foams may be of semi-rigid or rigid grades consisting of closed cells with characteristic low compression set and heat transmission [9]. Several types of conventional polyols are available in the market today mostly formulated from patented works of several investigators [11]. Their backbones are based mainly on ethylene oxide or propylene oxide or their derivatives or blends of the two (equation1). Various percentage compositions of these oxides have been reported ranging from molecular weight averages (450 – 30,000) depending on the end-use requirements of the final product. The primary hydroxyl groups for these polyols have been reported to range from 2 – 6,

depending on the type of glycols (glycerol, pentaerithritol, sorbitol etc) incorporated in the formulation of the prepolymer [8], [10-11].

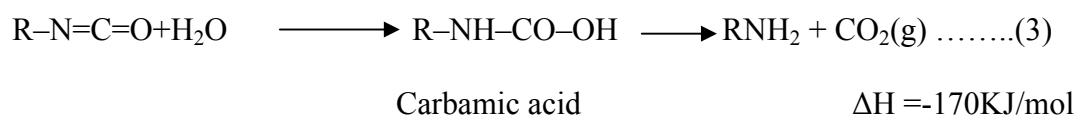


Polymer polyols are essentially the conventional polyols or their blends incorporated with a copolymer mainly styrene-acrylonitrile copolymer or styrene-acrylonitrile-vinylidene chloride terpolymer dispersed or dissolved within the conventional polyols to produce high density, high resilient, higher load bearing, non-shrinking, free rise foams compared to those from unmodified polyols. The compositions of these polymer polyols are dependant on the copolymerization of one or two ethylenically unsaturated monomers dissolved or dispersed in a conventional polyol to form a stable dispersion of polymer particles in the polyols [10-12].

Much have been reported and several patents are available on the development of polyurethane foam Chemistry and Technology since the end of World War II (1945). The concern here is that polyurethane foam formation is extremely an exothermic process, producing large amount of heat during the reaction process leading to trouble-shooting [13-14] Mekanjuola, (1999) and Ikeh, *et al*, (2007) in their separate reports indicated that manufacturers of polyurethane foams must understand the chemistry of the process, so that trouble-shooting is easily carried out with necessary formulation adjustment to compensate for local prevailing conditions. If the chemicals are too hot or cold, problems will be encountered as viscosity, density and chemical reactivity of polyols and TDI vary with temperature. Consequently, the monitoring and correlation of temperature rises with extent of reaction and physical properties are essential at optimal heat level. The first and chain propagating reaction is the basic reaction for the formation of urethane, that is polymerization reaction between polyol and TDI [15-16].



The second reaction is the gas –forming reaction, which liberates CO₂ and simultaneous formation of urea. This is the reaction between TDI and water. The first step in this reaction is the formation of an unstable carbamic acid, which decomposes to form a primary amine and carbon dioxide which serves as the primary blowing agent, with evolution of large amount of heat (equation 2). Therefore the overall reactions are extremely exothermic with temperature rises between 160-180⁰C [3].



Water serves as a primary blowing agent for polyurethane systems. Several researchers have reported the effect of water as a blowing agent in polyurethane foam manufacture and have observed that the large amount of heat liberated during this process has great effects on the physical and mechanical properties of foams manufactured and the causes of factory fire outbreaks [17]. Monitoring the effects of exotherms (equations 2 & 3) on foam densities , heights and compressions with variations in the amount of water added as blowing agent using conventional and polymer polyols is necessary as this has consequences on the comfort and mechanical properties during service life [18].

The aim of this research is to monitor and evaluate the effects of exotherms on some foam physical properties with variations in the amount of water added as blowing agent using conventional and polymer polyols.

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105 **Materials and Method**

106 **Samples/Industrial chemicals**

107 Conventional polyols (CPO)- polyether based

108 Polymer polyol (PPO)- polyether based with dispersed (grafted) copolymers such as ABS,
109 Acrylonitrile, Styrene etc.

110 Toluene diisocyanate (TDI 80:20) Bpt 250⁰C.

111 De-ionised water (blowing agent)

112 Stannous Octanoate- catalyst.

113 Dimethyl ethanolamine- catalyst.

114 Poly silicone fluid (polydimethyl siloxane) - surfactant.

115 Methylene dichloride- Flame retardant and auxiliary blowing agent

116 Calcium carbonate (CaCO₃)- filler (bulk and reinforcing filler).

117 All the chemicals, except water were gratefully obtained from Vitafoam Nig. Plc. Jos factory.

118 **Equipment/apparatus**

119 Transparent cylindrical plastic moulds(14cm- diameter by 20cm height) were locally assembled

120 Glasswares/vernier calipers, micrometer screw gauge, micro-syringes, thermometers, balances
121 etc were obtained in our laboratory.

122 **Determination of some physical properties of the polyols**

123 The refractive indices, specific densities and pH values of the polyols were determined by
124 conventional methods [19].

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130 **Determination of some chemical parameters of the polyols.**

131 **Determination of Acid value [20]**

132 **Method:** Three grams (3g) of the polyol was weighed into a 250cm³ conical flask and 40cm³ of

133 mixed solvent (2 parts benzene and 1 part ethanol (v/v) was added. This was dissolved by

rotating the flask and warming on a water bath. It was cooled and titrated rapidly with 0.1M ethanolic KOH until change of colour occurred. Similarly, a blank titration (without polyol sample) was concurrently run.

The acid values were calculated from the expression:

$$\text{Acid value} = \frac{56.1 \times M (a - b)}{w}$$

a = Volume of KOH (in cm³) for blank

b = Volume of KOH (in cm³) for polyol

M = molarity of KOH

w = weight of polyol sample (g)

Determination of Hydroxyl value of the polyols [11]

Acetylating agent: This was prepared by addition of sufficient amount of distilled pyridine to 25g of dry acetic anhydride in a 100cm³ volumetric flask and diluted to volume. The entire content was mixed by swirling the flask gently, and the mixture stored in a brown glass container, stoppered and placed in the fume cupboard.

Method: 1.5g of polyol was weighed into a 250cm³ flask and 5cm³ acetylating agent was added using a pipette. The mixture was shakened gently to ensure uniformity. A blank was also prepared by measuring an equal amount of the acetylating agent (5cm³) into a second 250cm³ flask. The two flasks were place in a steam bath and heated for one hour under reflux. They were allowed to cool and water was added through the condenser. The mixtures were shakened to ensure turbidity was absent. The mixtures were allowed to cool and the flasks and condenser washed down with 5cm³ ethanol to neutralize. Each mixture was then titrated with 0.5M ethanolic potassium hydroxide solution and 3drops of 1% ethanolic phenolphthalein indicator

added until a faint pink colour was reached. The hydroxyl value was calculated from the expression;

$$\text{Hydroxyl value} = \frac{56.1 \times M (v_2 - v_1)}{w} + A$$

Where: A = Acid value as calculated above. M = Molarity of KOH solution

v_2 = Volume of KOH solution used for blank

v_1 = Volume of KOH solution used for oil

w = weight of Polyol

Table 1a: Standard Formulations for Preparing CPO and PPO based foams for monitoring of exotherms

	Pbv (cm³)	Pbw(g)
CPO or PPO	10.00	10.40
TDI	4.50	5.54
Water	0.10	0.10
D.M.E.A	0.02	0.02
Silicone	0.10	0.10
Stannous Octonate	0.04	0.05
Methylene Chloride	1	1.5
Calcium Carbonate	--	2.10

Where Pbv and Pbw = Part by volume and part by weight respectively

Table 1b: Summary of Formulations for Exotherm Studies

Sample Identity for CPO	Formulations	Sample Identity for PPO
A ₁ ---A ₁₀	In the presence of CaCO ₃ and MC	X ₁ ---X ₁₀
B ₁ ---B ₁₀	In the presence of CaCO ₃ only	Y ₁ ---Y ₁₀
C ₁ ---C ₁₀	In the presence of MC only	W ₁ ---W ₁₀
D ₁ ---D ₁₀	In the absence of CaCO ₃ and MC	Z ₁ ---Z ₁₀

Monitoring of Exotherms in Foam Formulation [16]

Standard formulations were prepared from recipes as shown on Tables 1a and 1b for CPO and PPO. Transparent air-tight cylindrical plastic moulds (flasks) of 14cm-diameter and 20cm heights were assembled. With a motorized screw, two holes were drilled through the lid of each mould to insert a thermometer and a flighted glass stirrer to fit. The thermometer was inserted in such a way that the lower bulb goes three-quarters into the mould to enable easy contact with rising foam mix. The one-shot technique was employed whereby all the “ingredients” were added stepwisely using syringes and mixed thoroughly for 10mins before the Toluene Diisocyanate (TDI) was added and stirred to cream. The room temperature and the temperature at cream, and final temperature at rise were taken and recorded. Furthermore, the difference between room temperature and rise temperature is taken as ΔT_1 and between cream and rise as ΔT_2 . Various formulations were then carried out with water at 0.1, 0.2, 0.3, 0.4, 0.5, 0.6, 0.7, 0.8,

0.9, 1.0cm³ in that order and temperatures recorded, with some formulations carried out in the presence or absence CaCO₃ and MC as shown on Table 1b. The densities of the foams from the formulations are calculated as follows;

$$\text{Density} = \frac{\text{weight of foam (g)}}{\text{volume of foam (cm}^3\text{)}} = \frac{\text{weight of foam}}{\pi r^2 h}$$

Volume of foam = $\pi r^2 h$: r = radius of cylinder (cm), h = foam height (cm), $\pi = 22/7$,

Determination of % Compression Set [21]

The test pieces were cut into dimensions. The samples were placed in the compression device which consisted of two flat plates with parallel to each other was ensured and the space between plates was adjusted to the required deflected height. The test piece was compressed by 50%, maintained at ambient temperature of 27°C for 22 hours. At the end of the 22 hours, the test piece was removed from the device and placed on a wooden surface since wood is of low thermal conductivity. The test piece was allowed to recover for 30minutes after which its thickness was measured as (Tr)

$$\% \text{ Compression Set} = \frac{T_o - Tr}{T_o} \times 100$$

Where T_o = initial thickness, Tr = thickness after removal of load after 30 mins.

RESULTS AND DISCUSSION

The results of some physical and chemical properties of the polyols are presented on Table 2 while the results of the exotherms for the various foam formulations are shown on Tables 3-10.

Table 2 gives some physical and chemical properties of the polyols. It indicates that the Polymer polyol (PPO) is more viscous, acidic, with higher boiling point, density and invariably higher molecular weight than conventional polyol (CPO). However CPO has a higher hydroxyl number than PPO because it has more free primary OH-groups than PPO. The reason is that PPO as grafted co-polymer has a higher molecular weight with less primary hydroxyl groups than CPO. This implies that PPO would form more rigid and denser foams than CPO [8], [10], [22-23].

Table 2: Some physical and chemical Properties of the Conventional and polymer polyols

Properties	Conventional Polyol(CPO)	Polymeric Polyol (PPO)
Colour (visual)	Clear	milk white
State	Liquid	Liquid
Relative viscosity at 25⁰C	10.25	11.66
Solubility in alcohol	Complete	Complete
Boiling point (⁰C)	145	182
Refractive index	1.47	---
Density (g/cm³)	1.01	1.04
pH	10.25	5.52
Hydroxyl value (mgKOH/g)	187.20	169.20
Acid value (mg KOH/g)	1.16	2.23

Tables 3-10 present the results of co-relations of exotherms and blow indices with physical properties in the presence or absence of CaCO₃ filler or methylene

dichloride (MC), an auxiliary blowing agent and fire retardant. The results show the effects of variations of volume of water (a primary blowing agent) as blow indices in the presence or absence of MC (a prohibited volatile organic compound (VOC) and one of the CFCs banned by EU and the Federal Government of Nigeria [24-27], yet the foam industry in Nigeria still uses) or CaCO_3 a conventional filler, on exotherms, foam densities, heights and % compression. The results generally show that % compression and height increase with increase in blow index as exotherm also increases while density decreases. Generally it was observed that CPO foam formulations generate more heat (exotherms) compared to the ones formulated with PPO. This can be explained on the fact that CPO system has higher concentration of free primary OH-groups making it more reactive than PPO which has higher concentration of secondary OH-groups making less reactive comparatively [16], [23]. Furthermore, because of viscosity build up at gelation, the dissipation of heat at gelation stage for the CPO system is easily achieved since it is less viscous than PPO, leading to higher exotherms compared to PPO which does it slowly. Again, PPO cross-link density is higher than that of CPO leading to a corresponding gradual increase in temperature with increase in volume of water as blowing agent. Consequently, PPO foams have higher densities than CPO foams at increasing volume of water and also % compression [8]

Comparing and contrasting parameters on Tables 3 and 4, exotherms A_1 - A_{10} CPO versus X_1 - X_{10} PPO formulations respectively in presence of CaCO_3 and methylene chloride (MC) show, ΔT_1 (38-50⁰C); ΔT_2 (37-39.5⁰C) and ΔT_1 (32-42⁰C); ΔT_2 (29-36⁰C) in that order . These gradual increases are expected with increasing volume of water from 0.1-1.0 cm³ as more heat is liberated. This is expected as the volume of water increases, more CO_2 is generated leading to inflation of the foams, with increase in exotherms (equation 2 and 3). Furthermore ΔT_1 and ΔT_2

values for CPO formulations are correspondingly higher than ΔT_1 and ΔT_2 for PPO. However, X_1 - X_{10} PPO foam densities are higher than those of A_1 - A_{10} , CPO due to composition differentials earlier mentioned and the same reason goes for % compression which increases linearly with blow index and height for both CPO and PPO [17]

Table 3: Results of variations of exotherms and physical properties with blow index for Conventional Polyol (CPO) foam formulations in the presence of MC and CaCO_3 .

Foam identity	Blow index	Temperatures ($^{\circ}\text{C}$)					Physical properties		
		Room (t_o)	Cream (t_c)	Rise (t_f)	ΔT_1	ΔT_2	Density (g/cm^3)	Height (cm)	Compression (%)
A_1	0.16	26.00	27.00	64.00	38.00	37.00	22.90	8.00	2.90
A_2	0.17	26.00	27.50	65.50	39.50	38.00	22.50	8.50	3.23
A_3	0.18	26.00	28.00	69.00	41.00	39.00	21.90	8.90	3.27
A_4	0.19	26.00	29.00	70.50	43.00	40.00	21.70	9.90	3.74
A_5	0.20	26.00	31.00	71.80	44.50	39.50	21.43	9.95	4.09
A_6	0.21	26.00	31.50	72.00	45.80	40.30	21.05	10.00	4.27
A_7	0.22	26.00	32.50	73.50	46.00	39.50	20.85	10.48	4.65
A_8	0.23	26.00	33.00	75.00	47.50	40.30	20.60	10.85	5.09
A_9	0.24	26.00	36.50	75.00	49.00	40.00	20.30	10.90	5.73
A_{10}	0.25	26.00	36.50	76.00	50.00	39.50	19.70	11.05	6.18

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265 Where $\Delta T_1 = t_f - t_o$; $\Delta T_2 = t_f - t_c$

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268 **Table 4: Results of variations of exotherms and physical properties with blow index for**
 269 **Polymer Polyol (PPO) foam formulations in the presence of MC and CaCO₃.**

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Foam identity Sample code	Blow index	Physical properties							
		Temperatures (°C)			ΔT_1	ΔT_2	Density (g/cm ³)	Height (cm)	Compre ssion (%)
		Room (t _o)	Cream (t _c)	Rise (t _f)					
X ₁	0.16	26.00	29.00	58.00	32.00	29.00	23.60	7.20	2.50
X ₂	0.17	26.00	29.50	59.00	33.00	29.50	23.20	7.35	2.72
X ₃	0.18	26.00	30.00	62.50	36.00	32.00	23.00	8.40	3.00
X ₄	0.19	26.00	30.50	62.50	36.50	32.50	22.96	8.58	3.28
X ₅	0.20	26.00	31.00	63.80	37.50	32.80	22.95	8.60	3.94
X ₆	0.21	26.00	32.00	63.50	38.80	32.50	22.90	8.61	4.70
X ₇	0.22	26.00	33.00	64.00	38.00	31.50	22.50	9.65	5.63
X ₈	0.23	26.00	33.50	65.00	39.50	31.50	22.42	9.70	5.77
X ₉	0.24	26.00	33.50	67.00	41.50	32.00	22.35	10.72	6.01
X ₁₀	0.25	26.00	36.00	68.50	42.50	36.50	22.30	10.80	6.07

271 Where $\Delta T_1 = t_f - t_o$; $\Delta T_2 = t_f - t_c$

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Comparing and contrasting parameters on Tables 5 and 6, exotherms for B₁-B₁₀ CPO versus Y₁-Y₁₀ PPO formulations respectively in the presence of CaCO₃ only, shows; ΔT_1 (33-52⁰C); ΔT_2 (32-42.5⁰C) and ΔT_1 (17-28⁰C); ΔT_2 (14-18.8⁰C) in that order. Again, the same trend is observed as indicated on Tables 3 and 4. However, it was observed that the absence of MC has little or no effects on the exotherms ΔT_1 and ΔT_2 CPO formulations on Table 5 compared to the corresponding ΔT_1 and ΔT_2 CPO exotherms displayed on Table 3. On the other hand, the absence of MC has drastic effects on the exotherms, ΔT_1 and ΔT_2 PPO formulations on Table 6 compared to ΔT_1 and ΔT_2 PPO shown on Table 4 with ΔT_1 and ΔT_2 on Table 6 displaying much lower values than the corresponding ΔT_1 and ΔT_2 on Table 4. However, the corresponding densities are in reverse order. The reason for this phenomenon may be explained on the fact that the heat generated in the process is absorbed more or less by CaCO₃ present [15] [19]. Again compression (%) increases linearly with increases in blow indices and heights for both CPO and PPO samples.

298 **Table 5: Results of variations of exotherms and physical properties with blow index for**
 299 **CPO foam formulations in the presence of CaCO₃ only.**

Foam identity	Blow index	Physical properties							
		Temperatures (oC)			ΔT_1	ΔT_2	Density (g/cm ³)	Height (cm)	Compression (%)
Sample code		Room (t _o)	Cream (t _c)	Rise (t _f)					
B ₁	0.16	26.00	26.90	59.00	33.00	32.10	23.40	7.20	2.24
B ₂	0.17	26.00	27.00	62.00	36.00	35.00	23.00	7.40	2.26
B ₃	0.18	26.00	28.00	64.00	38.00	36.00	22.80	7.80	3.10
B ₄	0.19	26.00	30.00	65.50	39.50	35.50	22.45	7.89	3.43
B ₅	0.20	26.00	31.00	67.00	41.00	36.20	22.01	8.80	3.86
B ₆	0.21	26.00	32.80	69.00	43.00	36.50	21.90	8.85	3.91
B ₇	0.22	26.00	34.00	70.50	44.50	36.50	21.70	8.90	4.72
B ₈	0.23	26.00	34.50	71.00	45.00	36.50	21.30	9.07	4.98
B ₉	0.24	26.00	35.00	73.50	47.50	38.50	21.00	9.50	5.59
B ₁₀	0.25	26.00	36.00	78.50	52.50	42.50	20.85	10.20	6.13

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301 Where $\Delta T_1 = t_f - t_o$; $\Delta T_2 = t_f - t_c$

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Table 6: Results of variations of exotherms and physical properties with blow index for PPO foam formulations in the presence of CaCO₃ only.

Foam identity Sample code	Blow index	Temperatures (°C)				Physical properties			
		Room t _o	Cream t _c	Rise (t _f)	ΔT ₁	ΔT ₂	Density (g/cm ³)	Height (cm)	Compression (%)
Y ₁	0.16	26.00	29.00	43.00	17.00	14.00	25.80	5.50	1.11
Y ₂	0.17	26.00	30.00	43.50	17.50	13.50	25.75	6.00	1.18
Y ₃	0.18	26.00	30.50	45.00	19.00	14.50	25.00	7.20	2.01
Y ₄	0.19	26.00	31.00	47.50	21.00	16.00	24.95	8.00	2.45
Y ₅	0.20	26.00	31.00	47.50	21.50	16.50	24.70	8.50	2.87
Y ₆	0.21	26.00	33.00	48.00	22.00	16.00	24.40	8.90	3.11
Y ₇	0.22	26.00	34.00	49.00	23.00	17.00	24.10	9.00	3.75
Y ₈	0.23	26.00	34.50	51.00	25.00	16.50	24.00	9.50	4.55
Y ₉	0.24	26.00	35.00	53.00	27.00	18.00	23.80	10.00	4.98
Y ₁₀	0.25	26.00	36.00	54.80	28.80	18.80	23.45	10.10	5.07

Where $\Delta T_1 = t_f - t_o$; $\Delta T_2 = t_f - t_c$

Comparing and contrasting parameters on Tables 7 and 8, exotherms for C₁-C₁₀ CPO versus W₁-W₁₀ PPO formulations respectively in the presence of MC only, show, ΔT_1 (35-50⁰C); ΔT_2 (33-41.3⁰C) and ΔT_1 (22-38.8⁰C); ΔT_2 (20-27⁰C) in that order. Again it was observed that the presence of MC has little or no effects on CPO formulations compared to the results on Table 3. However, the presence of MC only, has a drastic effect on the exotherms on PPO formulations; as ΔT_1 and ΔT_2 PPO exotherms on Table 8 are much lower than the corresponding ΔT_1 and ΔT_2 on Table 4. This means that MC suppresses heat evolved easily in PPO formulations than in CPO which is more reactive [11]. Compression (%) increases linearly with increases in blow indices and heights for both CPO and PPO samples.

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340 **Table 7: Results of variation of exotherms and physical properties with blow index for**
341 **CPO foam formulations in the presence of MC only.**

Foam identity Sample code	Blow index	Room t_0	Cream t_c	Temperatures ($^{\circ}\text{C}$)		ΔT_1	ΔT_2	Physical properties		
				Rise (t_f)				Density (g/cm^3)	Height (cm)	Compression (%)
C ₁	0.16	26.00	28.00	61.00		35.00	33.00	19.80	8.70	3.17
C ₂	0.17	26.00	28.50	62.50		36.50	34.00	19.45	8.80	3.20
C ₃	0.18	26.00	29.00	63.00		37.00	34.00	18.00	8.94	4.03
C ₄	0.19	26.00	30.00	65.00		39.00	35.00	18.90	9.00	4.29
C ₅	0.20	26.00	31.50	67.20		41.20	35.70	18.71	9.90	4.67
C ₆	0.21	26.00	31.90	68.90		42.90	37.00	18.46	9.58	4.98
C ₇	0.22	26.00	32.50	69.00		43.00	36.50	17.13	9.95	5.15
C ₈	0.23	26.00	33.00	72.00		46.00	39.00	17.01	10.20	6.79
C ₉	0.24	26.00	34.50	73.00		47.00	38.50	18.82	10.82	7.77
C ₁₀	0.25	26.00	35.20	76.50		50.00	41.30	18.51	11.00	8.53

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343 Where $\Delta T_1 = t_f - t_0$; $\Delta T_2 = t_f - t_c$

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Table 8: Results of variations of exotherms and physical properties with blow index for PPO foam formulations in the presence of MC only.

Foam identity	Blow index	Temperatures ($^{\circ}\text{C}$)					Physical properties		
		Room t_o	Cream t_c	Rise (t_f)	ΔT_1	ΔT_2	Density (g/cm^3)	Height (cm)	Compression (%)
W ₁	0.16	26.00	28.00	48.00	22.00	20.00	15.50	9.90	3.01
W ₂	0.17	26.00	28.00	48.50	22.50	20.50	15.00	10.50	3.17
W ₃	0.18	26.00	28.00	50.00	24.00	22.00	14.90	10.90	3.39
W ₄	0.19	26.00	28.00	52.00	26.00	24.00	14.24	11.20	3.72
W ₅	0.20	26.00	30.00	55.50	29.50	25.50	14.00	11.80	4.12
W ₆	0.21	26.00	31.50	57.00	31.00	25.50	13.80	11.80	5.16
W ₇	0.22	26.00	33.00	58.50	32.50	25.00	13.40	12.10	5.56
W ₈	0.23	26.00	34.00	61.00	35.00	27.00	13.00	12.50	6.78
W ₉	0.24	26.00	34.50	63.00	37.00	28.50	12.50	12.60	7.46
W ₁₀	0.25	26.00	37.00	64.80	38.80	27.80	12.00	12.80	8.20

Where $\Delta T_1 = t_f - t_o$; $\Delta T_2 = t_f - t_c$

Comparing and contrasting parameters on Tables 9 and 10 exotherms for D₁-D₁₀ CPO versus Z₁-Z₁₀ PPO formulations respectively in the absence of CaCO₃ and MC indicates that ΔT_1 (34-49.2⁰C); ΔT_2 (33-41.2⁰C) and ΔT_1 (36-54⁰C); ΔT_2 (35-47.1⁰C) in that order. It was observed again that both CaCO₃ and MC has no effects on exotherms of CPO formulations compared to CPO exotherms on Table 3 with added CaCO₃ and MC, even though corresponding densities are higher compared to those on Table 9 because CaCO₃ as a filler increases density. On the other hand, the absence of CaCO₃ and MC have tremendous effect on the PPO exotherms as ΔT_1 and ΔT_2 PPO exotherms on Table 10 in the absence of CaCO₃ and MC are higher than the corresponding exotherms on Table 4 in the presence of CaCO₃ and MC. Similarly, PPO formulation exotherms on Tables 4, 6 and 8 in the presence of MC only, have much lower values compared to the corresponding exotherms on Table 10. This suggests that the effects of MC as a fire (flame) retardant on PPO formulations are more pronounced, but its effects are overshadowed by the high reactivity of the free-primary OH-groups in CPO. In any case, it blends (interacts) favorably with PPO [14]. Compression (%) increases linearly with increases in blow indices and heights for both CPO and PPO samples while density shows an inverse relationship with these parameters.

Table 9: Results of variations of exotherms and physical properties with blow index for CPO foam formulations in the absence of MC and CaCO₃.

Foam identity Sample code	Blow index	Temperatures (°C)				Physical properties			
		Room t _o	Cream t _c	Rise (t _f)	ΔT ₁	ΔT ₂	Density (g/cm ³)	Height (cm)	Compr ession (%)
D ₁	0.16	26.00	27.00	60.00	34.00	33.00	23.40	5.50	2.63
D ₂	0.17	26.00	28.00	61.50	35.50	33.50	23.01	6.10	2.91
D ₃	0.18	26.00	29.00	63.00	37.00	34.00	22.50	6.70	3.24
D ₄	0.19	26.00	29.50	63.50	37.50	34.00	22.04	6.95	3.97
D ₅	0.20	26.00	30.00	64.00	38.00	34.00	22.00	7.60	4.11
D ₆	0.21	26.00	30.20	66.00	40.00	35.80	21.35	7.51	4.97
D ₇	0.22	26.00	31.50	68.00	42.00	36.50	21.35	7.75	5.73
D ₈	0.23	26.00	31.00	69.80	43.80	37.80	21.08	8.00	5.97
D ₉	0.24	26.00	33.00	72.00	46.00	39.00	20.90	8.70	6.29
D ₁₀	0.25	26.00	34.00	75.20	49.20	41.20	20.75	8.85	7.61

Where $\Delta T_1 = t_f - t_o$; $\Delta T_2 = t_f - t_c$

387 **Table 10: Results of variations of exotherms and physical properties with blow index for**
 388 **PPO foam formulations in the absence of MC and CaCO₃.**

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Foam identity	Blow index	Temperatures (°C)					Physical properties		
		Room t _o	Cream t _c	Rise (t _f)	ΔT ₁	ΔT ₂	Density (g/cm ³)	Height (cm)	Compression (%)
Z ₁	0.16	26.00	27.00	62.00	36.00	35.00	21.00	4.90	2.97
Z ₂	0.17	26.00	27.50	63.70	37.70	36.20	20.90	5.50	3.03
Z ₃	0.18	26.00	28.50	66.00	40.00	37.50	20.50	5.90	3.18
Z ₄	0.19	26.00	28.50	68.50	42.50	39.50	20.00	6.20	3.47
Z ₅	0.20	26.00	29.00	70.00	44.00	41.00	19.80	6.80	4.01
Z ₆	0.21	26.00	30.50	71.00	45.00	40.50	19.20	7.00	4.22
Z ₇	0.22	26.00	31.00	73.50	47.50	42.50	19.00	7.10	4.67
Z ₈	0.23	26.00	32.00	74.00	48.00	42.00	18.70	7.50	5.99
Z ₉	0.24	26.00	32.50	78.00	52.00	45.50	18.50	8.00	6.38
Z ₁₀	0.25	26.00	32.90	80.00	54.00	47.10	18.10	8.80	7.88

390 Where ΔT₁ = t_f - t_o; ΔT₂ = t_f - t_c

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399 CONCLUSION

400 From this study, the following conclusion can also be made; the concentration of
 401 blowing agent invariably, relates to the physico-mechanical properties of foams hence
 402 monitoring of exotherms is very critical [28-29] . The production of polyurethane foams is an
 403 exothermic process, liberating large amount of heat which depends on the following factors; type
 404 and functionality of the polyol, and type and amount of blowing agent. The temperature-amount
 405 of water profile reveals the extent of heat liberated using CPO and PPO. The relative importance
 406 of both the blowing and the gelling steps can be assessed by their individual contribution to the
 407 overall exotherms. The CPO based formulations liberate more heat than the corresponding
 408 polymer polyol based formulations which relate to their compositions as earlier mentioned [ref
 409]Based on foam densities and compositions, the polymer polyol is more suitable for both high
 410 and low density rigid and flexible polyurethane foams while the conventional polyol on the other
 411 hand is suitable for very low density flexible polyurethane foams [30-31].

412 RECOMMENDATION

413 It is recommended that more research should be carried out to find ways of reducing exotherms
 414 in polyurethane foam manufacture. Appropriate mold design should be employed to minimized
 415 the adverse effect of high exotherms on the end products.

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