

Synthesis, spectral, magnetic and *in-vitro* anticancer properties of some Metal(II) complexes of 3-[2,4-dihydro-1H-inden-4-ylimino) methyl]naphthalen-2-ol.

ABSTRACT

Novel Cu(II), Zn(II) and Pd(II) complexes of the Schiff base, 3-[2,4-dihydro-1H-inden-4-ylimino)methyl]naphthalen-2-ol, have been prepared and characterized by elemental analysis, ¹HNMR, molar conductance, room temperature magnetic moments, infrared and electronic spectral measurements. The purity of the ligands and the metal complexes are confirmed by microanalyses and ¹HNMR. The IR spectra shows that the Pd(II) exhibits geometric isomerism, being in the trans isomeric form, and the ligand is bidentate coordinating via the azomethine N and naphthalenol O atoms. The Cu(II) and Pd(II) complexes assume a four-coordinate square planar geometry; while the Zn(II) complex is tetrahedral, as indicated by room temperature moment and electronic spectra measurements. The complexes are non-electrolytes in nitromethane. The *in-vitro* anticancer activities of the Schiff base and its Cu(II), Zn(II) and Pd(II) complexes against HT-29 (colon) carcinoma and MCF-7 (human breast) adenocarcinoma are evaluated. The results show that the HT-29 cells are generally more sensitive than the MCF-7 cells to the metal complexes; and the Zn(II) complex has the best broad-spectrum anticancer activity against both colon and breast carcinomas, with IC₅₀ values of 20.15 and 22.50 μM respectively. The activity of the Zn(II) complex against colon carcinoma is about a third of Cis-platin activity in the same assay.

Key words: anticancer, broad-spectrum; *in-vitro*, non-electrolyte, Schiff base.

1. INTRODUCTION

Aminoindane based derivatives exhibit a range of bioactivities such as bronchodilators (Heinzelmann *et al.*, 1948), anticancer, antimicrobial and anticonvulsant (Maruyama *et al.*, 2001; Rao and Reddy, 1990; Rogers *et al.*, 1989; Stewart *et al.*, 1996); and as analgesics and protease inhibitors (Borenstein *et al.*, 1984; Zon *et al.*, 2005). Furthermore, Cr(III), Ru(III) and Ti(IV) complexes of various aminoindanyl Schiff bases are used as catalysts in Diels-Alder reactions of α,β-unsaturated aldehydes and ring opening of mesoaziridines, in the oxidation of benzyl alcohol and cyclohexanol to benzaldehyde and cyclohexanone; and in the oxidation of alkyl aryl sulfides to the corresponding sulfoxides (Bryliakov *et al.*, 2007; Fernandez and Jacobsen., 1999; Gademann *et al.*, 2002). In addition, VO(IV) complexes of some symmetric 2-hydroxynaphthaldehyde Schiff bases have been found to be promising

anti-diabetic agents (Nejo *et al.*, 2009). In-depth literature search revealed that no information is available on the Schiff base, 3-[2,3-dihydro-1H-inden-4-ylimino)methyl] naphthalen-2-ol (derived from condensation of 2-hydroxy-1-naphthaldehyde and 4-aminoindane) and its Cu(II), Zn(II) and Pd(II) complexes (Abdullah and Badahdah, 2007; Blagus *et al.*, 2010; Derebe *et al.*, 2002; Kozlov and Gusak, 1999; Na *et al.*, 1988; Nejo *et al.*, 2010a, 2010b; Santos *et al.*, 2000; Tuncel and Selahattin, 2005). Therefore the aim of our study is primarily to synthesize and characterize the above named Schiff base and its metal (II) complexes, and evaluate their suitability as anticancer agents in-vitro. Secondly, the compounds with very good in-vitro anticancer activities will be used as treatments in-vivo in albino mice with various forms of cancer (this result will be presented in another paper). This study is a continuation of the research activities of our group in the search for lead compounds for drug development against various carcinomas (Osohole and Daramola 2011; Osohole *et al.*, 2010a, 2010b; Osohole, 2011).

2. MATERIAL AND METHODS

2.1 Chemicals

Reagent grade 2-hydroxy-1-naphthaldehyde, 4-aminoindane, triethylamine, Copper(II) nitrate hexahydrate, Palladium(II) chloride dihydrate, and Zinc(II) nitrate heptahydrate were purchased from Aldrich chemicals and were used as received.

2.2 Syntheses

2.2.1 Preparation of 3-[2,4-dihydro-1H-inden-4-ylimino)methyl]naphthalen-2-ol

The ligand was prepared by refluxing for 3 h, a homogeneous solution of 1.94 g (1.13×10^{-2} moles) 2-hydroxy-1-naphthaldehyde and 1.5 g (1.13×10^{-2} moles) of 4-aminoindane in 60 mL of hot ethanol, to which 4 drops of acetic acid were added. The orange products [Figure 1(A)] formed on cooling in ice were collected by suction, recrystallized from ethanol and dried over silica gel. The yield of the product was 2.59 g (80%).

Color(yellow); IR(KBr, cm^{-1}): ν_{OH} (3424s), $\nu_{\text{C=N}}$ (1620s 1580s 1540s); UV, λ_{Max} (ϵ): 24.74 (1×10^4), 31.25(1×10^5), 41.32(1×10^6); M.pt($^{\circ}\text{C}$), 158-160; formula mass (287.36); CHN Anal. calcd(found) for $\text{C}_{16}\text{H}_{14}\text{N}_2\text{O}_3$: C, 83.60(83.63); H, 5.96(5.80); N, 4.87(4.65); $^1\text{Hnmr}$ (300 MHz, CDCl_3 , δ in ppm): 15.76 (s, OH), 9.4 (s, HCN), 7.30 - 8.14(m, 6H, C_{10}H_6); 7.0-7.27 (m, 3H, {C5' C6' C7' H₃}); 3.07-3.12(t, 2H, C3'H₂); 2.98-3.05 (t, 2H, C1'H₂); 2.10-2.22 (q, 2H, C2'H₂).

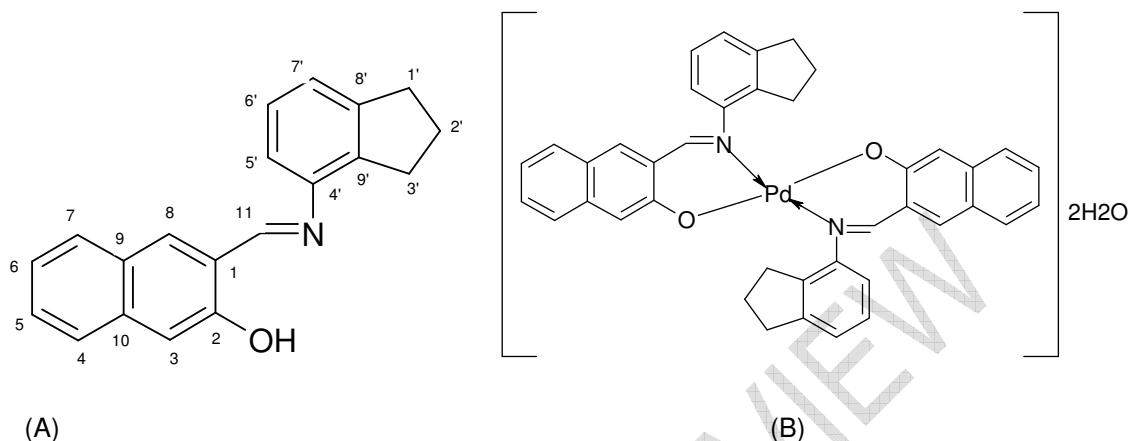


Fig. 1. The proposed structure for the ligand (A) and its Pd(II) complex (B).

2.2.2 Preparation of the metal complexes

The Cu(II) complex was prepared by refluxing for 3 h a homogeneous solution of 0.30 mmol (0.06 g) of Cu(II) nitrate hexahydrate and 0.60 mmol (0.17 g) of the Schiff base, to which 0.6 mmol (0.06 g) of triethylamine was added in 30 mL ethanol. The products precipitated as solids, which were filtered, washed with ethanol and dried over silica gel. The Pd(II) and the Zn(II) complexes were prepared from their chloride and nitrate salts using the same method.

The analytical data were as follow;

[CuL₂]³/₄H₂O: % yield 60 (0.12 g); Color(brown); IR(KBr, cm⁻¹): νOH (3500b), νC=N (1611s 1600s 1578s 1534m), νM-N (565s 521s) νM-O (498s 464s); VIS/UV, kK, λ_{Max} (ε): 14.87(200), 23.61(500), 31.20(1 × 10⁴), 38.0(1 × 10⁵), 40.98(1 × 10⁶); M.pt(°C), 208-210; formula mass (649.77); CHN Anal. calcd (found) for Cu(C₄₀H_{33.5}N₂O_{2.75}) C, 73.97(74.09); H, 5.20(4.77); N, 4.31(3.63); %Cu calcd (found)9.78(9.80); Λ_m, 10.0; μ_{eff}, 1.93.

[ZnL₂]7H₂O: % yield 40(0.09 g); Color(Red); IR(KBr, cm⁻¹): νOH (3500b), νC=N (1618s 1582s 1542s), νM-N (514s 502s) νM-O (498s 472s); VIS/UV, kK, λ_{Max} (ε): 21.10(3 × 10³), 31.20(1 × 10⁴), 41.32(1 × 10⁶); M.pt(°C), 230; formula mass (764.23); CHN Anal. calcd (found) for Zn(C₄₀H₄₆N₂O₉) C, 62.87(62.80); H, 6.07(4.58); N, 3.67(3.49); %Zn calcd (found) 8.51(8.48); Λ_m, 36.0; μ_{eff}, D; ¹Hnmr (300 MHz, CDCl₃, δ in ppm): 9.3 (s, HCN), 7.25-8.07 (m, 6H, C₁₀H₆); 7.02-7.24 (m, 3H, {C5' C6' C7' H₃}); 3.10-3.15(t, 2H, C3'H₂); 2.99-3.06 (t, 2H, C1'H₂); 2.13-2.23(q, 2H, C2'H₂).

[PdL₂]2H₂O: % yield 70(0.15 g); Color(Green); IR(KBr, cm⁻¹): νOH (3500b), νC=N (1607s 1579s), νM-N (576s 501s) νM-O (474s 428s); VIS/UV, kK, λ_{Max} (ε): 16.0(200), 23.41(3 × 10³), 30.48(1 × 10⁴), 36.76(1 × 10⁶), 40.0(1 × 10⁶); M.pt(°C), 280; formula mass (715.14); CHN Anal. calcd(found) for Pd(C₄₀H₃₆N₂O₄) C, 67.18(67.19); H, 5.07(4.46); N, 3.92(3.62); %Pd calcd(found)14.88(14.88); Λ_m, 29.0; μ_{eff}, D; ¹Hnmr (300 MHz, CDCl₃, δ in ppm): 8.58 (s, HCN), 7.27-8.04 (m, 6H, C₁₀H₆); 7.1-7.26 (m, 3H, {C5' C6' C7' H₃}); 3.17-3.23(t, 2H, C3'H₂); 3.01-3.06 (t, 2H, C1'H₂); 2.01-2.09 (q, 2H, C2'H₂).

2.3 Physical measurements

The electronic and infrared spectra were recorded on a Perkin-Elmer λ 25 and a Thermo Nicolet FTIR 200 spectrophotometer respectively. The ^1H nmr spectrum was recorded on a 300 MHz Bruker DRX-400 NMR instrument in CDCl_3 at 295K; and ^1H chemical shifts were referenced to the residual signals of the protons of CDCl_3 and were quoted in ppm. The elemental analyses for C, H and N were recorded on a Thermo Quest CE Instruments flash EA1112 analyser, while percentage copper, palladium and zinc were determined titrimetrically (Bassett *et al.*, 1978). Magnetic susceptibilities were measured on a Johnson Matthey magnetic balance at room temperature of 26°C , while diamagnetic corrections are calculated using Pascal's constants (Earnshaw, 1980). Electrolytic conductivities in nitromethane and melting points (uncorrected) were determined using a HANNA HI 991300 conductivity meter and Mel-Temp electro thermal machine respectively.

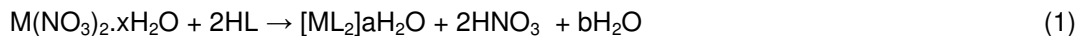
2.4 Anti-cancer studies

The MCF-7 (human breast adenocarcinoma) and HT-29 (colon carcinoma) cells were cultured and maintained in minimum essential medium (MEM) supplemented with 10% (V/V) of fetal calf serum (FCS) and 50 mg/L of gentamycin at the Institute of Medicinal and Pharmaceutical Chemistry, Technical University Braunschweig, Germany. A concentration of 100 mL of a cell suspension in a culture medium [7500 cells/mL (MCF-7) and 2500 cells/mL (HT-29)] were plated into each of 96 well plates and incubated for three days under culture conditions. After the addition of various concentrations of the test compounds, cells were incubated for another 96 h and 72 h respectively. The medium was then removed and the cells were fixed with 1% glutardialdehyde solution and stored under phosphate buffered saline (PBS) at 4°C . Cell biomass was determined by crystal violet staining, followed by the extraction of the bound dye with ethanol, and a photometric measurement at 590 nm. The test compounds were prepared fresh as stock solutions in DMF and diluted with the cell culture medium to the final assay concentrations (0.1% V/V DMF). Cis-platin was used as the reference drug. The IC_{50} value was determined as the concentration causing 50% inhibition of cell proliferation and calculated as the mean of at least two independent experiments (Rubner *et al.*, 2010).

3. RESULTS AND DISCUSSION

3.1 Analytical data and conductance measurements.

The reactions for the formation of the metal(II) Schiff base complexes are presented in equations 1 and 2 respectively. The compounds were formed in good yields of 60-70% with the exception of the Zn(II) complex, which had a yield of 40%.



{When $\text{M(II)} = \text{Cu}$, $x = 6$, $a = 0.75$, $b = 5.25$; Zn(II) , $x = 7$, $a = 7$, $b = 0$ }



The water molecules associated with the complex formation are outside the coordination sphere. This is because heating the hydrated complexes in a programmed oven in the temperature range $80\text{--}100^\circ\text{C}$ for 1-3 hours leads to dehydration, with corresponding weight loss of 0.75, 2 and 7 water molecules respectively. It has been documented that the water of

coordination is usually eliminated in the temperature range 120–250°C (Nejo *et al.*, 2010; Refat *et al.*, 2010). The structural formula of the ligand is provided in Figure 1(A); and its formation is confirmed by microanalysis and ¹HNMR. The ligand melted in the range 158–160°C, while the complexes melted/decomposed in the range 208–280°C, confirming coordination. Single crystal of the complexes could not be isolated from any solution. Hence, analytical, magnetic and spectroscopic data have been used to propose possible structures, and the complexes are reported here for the first time.

The metal(II) complexes are non-electrolytes in nitromethane as shown by their molar conductivities (Λ_m) of 10.0–36.0 ohm⁻¹cm²mol⁻¹, since a value of 60–118 ohm⁻¹ cm²mol⁻¹ is expected for 1:1 electrolyte (Geary, 1971).

3.2 Electronic Spectra and magnetic moments.

The ultraviolet spectra of the compounds in chloroform are characterized by three peaks between 30.48–31.25, 36.76–39.00 and 40.82–42.00 kK with molar absorptivities of 10⁴–10⁶ M⁻¹ cm⁻¹, assigned to π - π^* and charge transfer transitions (of various origins), respectively (Nejo *et al.*, 2009; Nejo *et al.*, 2010b). The molar absorptivities of the complexes in the visible region are in the range 10²–10³ M⁻¹cm⁻¹ ruling out octahedral geometry, since octahedral complexes have molar absorptivities in the range 1–50 M⁻¹cm⁻¹ (Nejo *et al.*, 2010a).

Regular tetrahedral compounds of copper(II) usually have a single broad band of $\approx 10^2$ molar intensity below 10.0 kK, while regular octahedral complexes have a single broad band above 10.0 kK, and square-planar complexes usually have two bands between 10.0–20.0 kK with molar absorptivities of 10²–10³ M⁻¹ cm⁻¹ (Tuncel and Selahattin, 2005). The Cu(II) complex studied, displays two absorption bands at 14.87 and 23.61 kK, and is consequently square planar with bands assigned to the $^2B_{1g} \rightarrow ^2A_{1g}$ and $^2B_{1g} \rightarrow ^2E_{1g}$ transitions. An observed moment of 1.93 B.M indicates that it is monomeric and complimentary of square-planar geometry (Nejo *et al.*, 2010b).

The Pd(II) complex exhibits two absorption bands at 16.00 and 23.41 kK, typical of square planar geometry and are assigned to the $^1A_{1g} \rightarrow ^1B_{1g}$ and $^1A_{1g} \rightarrow ^1E_{2g}$ transitions. It is expectedly diamagnetic (Osowole and Daramola, 2011; Osowole, 2011). As expected, the Zn(II) complex has no d-d band but a M→L CT band at 21.10 kK, and its diamagnetism is corroborative of a 4-coordinate tetrahedral geometry (Rao and Reddy, 1990).

3.3 Infrared and ¹Hnmr Spectra

The assignments of the infrared bands were made by comparing the spectra of the compounds with reported literature on similar systems (Abdullah and Badahdah, 2007; Blagus *et al.*, 2010). The uncoordinated C=N stretching vibrations occur as three bands expectedly at 1620–1540 cm⁻¹ in the ligand (Derebe *et al.*, 2002; Na *et al.*, 1988). Furthermore, they are bathochromic shifted to 1618–1534 cm⁻¹ on coordination to the metal ions as three to four bands, except for the Pd(II) complex which has two bands, indicative of the involvement of the N donor atom of C=N in coordination. The possession of two $\nu(C=N)$ bands by the Pd(II) complex indicates that it exhibits geometric isomerism, and it is in the trans isomeric-isomeric {Figure 1(B)} form (Nejo *et al.*, 2010a). The strong band at 3424 cm⁻¹ in the ligand is assigned as νOH stretching frequency, and it is conspicuously absent in the spectra of the complexes, indicative of the deprotonation and the involvement of the naphthalenol O in chelation. The new broad band at 3500 cm⁻¹, in the complexes is assigned to $\nu(OH)$ crystallization water. The presence of bands due to $\nu(M-O)$ and $\nu(M-N)$ at 498–428 and 576–501 cm⁻¹ respectively is further evidence of coordination. These

bands are absent in the ligand (Blagus *et al.*, 2010; Derebe *et al.*, 2002; Tuncel and Selahattin, 2005).

The phenolic proton in HL is observed at 15.76 ppm, while the imine proton is seen as a singlet at 9.40 ppm. The naphthalene protons resonate as a multiplet at 7.30-8.14 ppm while the protons at C5', C6' and C7' in the indane ring are observed as a multiplet at 7.0-7.27 ppm. The 2H at C1' is seen as a triplet at 2.98-3.05 ppm, while those at C2' resonate as a quintet at 2.10-2.22 ppm. The 2H at C3' resonate as a triplet at 3.07-3.12 ppm. Furthermore, the [ZnL₂] spectrum shows the absence of the phenolic proton at 15.76 ppm, which confirms coordination through the phenolic O atom. The imine proton is seen as a singlet at 9.30 ppm, the naphthalenol protons resonate as a multiplet at 7.25-8.07 ppm, while the protons on C5', C6' and C7' in the indane ring resonate as a multiplet at 7.02-7.24 ppm. All these are downshifted. On the contrary, the 2H on C1', C2' and C3' are seen as a triplet, a quintet and a triplet respectively at 2.99-3.06 ppm, 2.13-2.23 ppm and 3.10-3.15 ppm; and they are all up shifted. These shifts are indicative of coordination through the imine N atom (Osohole *et al.*, 2010b). Similarly, the spectrum of [PdL₂] shows the absence of the phenolic proton at 15.76 ppm, which confirms coordination through the phenolic O atom. The imine proton is seen as a singlet at 8.58 ppm. The naphthalenol protons and those on C5', C6' and C7' in the indane ring resonate as multiplets at 7.24-8.04 and 7.10-7.26 ppm respectively and are downshifted. The 2H on C1' and C3' are up shifted and seen as a triplet each at 3.01-3.06 ppm and 3.17-3.23 ppm respectively; while the 2H at C2' are seen as a quintet at 2.01-2.09 ppm and are down shifted. These shifts are indicative of coordination through the imine N atom (Osohole *et al.*, 2011).

3. 4 Anticancer Activity

The anticancer activities are presented in Table 1. The HT-29 cells are generally more sensitive to the compounds than MCF-7 cells. The metal free ligand is more active against HT-29 cells than the Cu(II) and Pd(II) complexes, but is less active than the Zn(II) complex. The increasing order of activity is [CuL₂] < [PdL₂] < HL < [ZnL₂]. In contrast, coordination enhances the activity of the metal complexes against MCF-7 cells, and the increasing order of activity is HL < [PdL₂] < [CuL₂] < [ZnL₂]. In both cases, the Zn(II) complex has the best in-vitro anticancer activity against HT-29 and MCF-7 cells. Furthermore, the enhanced activity of [ZnL₂] (IC₅₀ value of 20.15 µM) which is about a third of Cis-platin activity (IC₅₀ value of 7.0 µM) in the same assay against HT-29 cells, is of interest in the development of tumor therapeutics.

Table 1 *In-vitro* anticancer activities of the ligand and its Cu(II), Zn(II) and Pd(II) Complexes

Compound	IC ₅₀ [µM] Human breast adenocarcinoma (MCF-7)	IC ₅₀ [µM] Colon carcinoma (HT-29)
CDDP	2.0	7.0
HL	52.50±0.06	26.28±0.01
[CuL ₂] ³ / ₄ H ₂ O	29.89±0.09	55.98±0.35
[ZnL ₂]7H ₂ O	22.50±0.05	20.15±0.01
[PdL ₂] 2H ₂ O	49.95±0.01	43.12±0.25

CDDP = *Cis-platin*

4. CONCLUSION

The Schiff base, 3-[2,4-dihydro-1H-inden-4-ylimino)methyl]naphthalen-2-ol, coordinates to the Cu(II), Zn(II) and Pd(II) ions using the azomethine N and naphthalenol O atoms. The complexes assume a 4-coordinate tetrahedral/square-planar geometry, as corroborated by room temperature magnetic moments, infrared and electronic spectral measurements. The Pd(II) complex exhibits geometric isomerism being in the trans isomeric form. The in-vitro anticancer activities of the Schiff base and its Cu(II), Zn(II) and Pd(II) complexes against HT-29 (colon) carcinoma and MCF-7 (human breast) adenocarcinoma show that the HT 29 cells are generally more sensitive than the MCF-7 cells to the metal complexes, with the Zn(II) complex having the best broad-spectrum anticancer activity against both colon and breast carcinomas with IC₅₀ values of 20.15 and 22.50 μ M respectively.

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COMPETING INTERESTS

Authors have declared that no competing interests exist.

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