



SDI Review Form 1.6

Journal Name:	International Research Journal of Pure and Applied Chemistry
Manuscript Number:	2014_IRJPAC_10977
Title of the Manuscript:	Investigations of ferrous thiamine complex by cyclic voltammetry
Type of the Article	Original Research Article

General guideline for Peer Review process:

This journal's peer review policy states that **NO** manuscript should be rejected only on the basis of '**lack of Novelty**', provided the manuscript is scientifically robust and technically sound.

To know the complete guideline for Peer Review process, reviewers are requested to visit this link:

(<http://www.sciencedomain.org/page.php?id=sdi-general-editorial-policy#Peer-Review-Guideline>)



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PART 1: Review Comments

	Reviewer's comment	Author's comment (if agreed with reviewer, correct the manuscript and highlight that part in the manuscript. It is mandatory that authors should write his/her feedback here)
Compulsory REVISION comments	The article is incomplete. And is not deeply expressed.(And it was	
Minor REVISION comments	1- Figure 2: it is better to be corrected. Please check 2- Anodic and cathodic current should be shown correctly., For example, 3- The authors explain the increase and decrease in pH. Why is acidic selected? 4- which PH is selected? 5- Page 7 of 8: mechanisms should be written correctly.. Please check 6- Figure 5: it is better to be corrected., Please 7- Vltamvgram process cycle should be fully written. 8- Factors affecting Cape of the current density should be written 9- Ratio of the anode to the cathode current has not been written. 10- Vltamvgram cyclic of figure 3 should be explored More. Please check 11- The article is incomplete. And is not deeply expressed.(And it was	1. I had corrected the fig 2 and correctly mention cathodic and anodic current. 2. Acidic pH is selected because complex of ferrous thiamine is formed at about 6 pH. Other pH also used but they did not given proper results. 3. Mechanism is written again as “ $[AFe] \xrightleftharpoons{k_o \text{ fast}} A^+Fe + e^-$ $A^+Fe + e^- \xrightarrow{k_1 \text{ slow}} A^+ + Fe$ $A^+ + e^- \rightarrow A$ $A + Fe \xrightarrow{k_1} [AFe]$ In case of A₂Fe



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		$AFe_2 \xrightleftharpoons[k_0 \text{ fast}]{k_o \text{ fast}} A^+ Fe_2 + e^-$ $AFe_2 \xrightleftharpoons[k_0 \text{ fast}]{k_o \text{ fast}} A^+ Fe_2 + e^- \xrightarrow[k_{-1} \text{ slow}]{k_1} A^+ Fe + Fe + e^- \rightarrow AFe_2$ $A^+ Fe \xrightarrow{k_{-2}} A^+ + Fe$ $A^+ + e^- \rightarrow A$ $A + 2Fe \xrightarrow{k_1} [AFe_2]$ Randles Sevcik equation is used which shows the reversible reaction and under diffusion control system.
Optional/General comments	Figure 2: it is better to be corrected. Please check Figure 5: it is better to be corrected., Please	Fig 2 is corrected again Fig 5 is corrected again and tried they explain the comments very well.